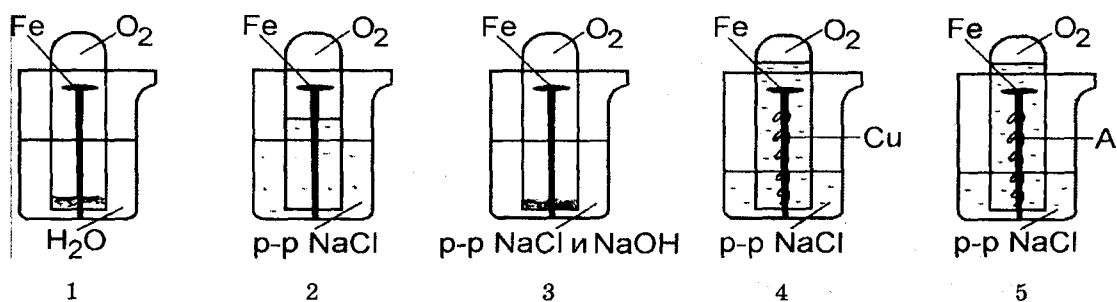


1.



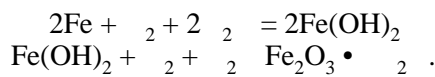
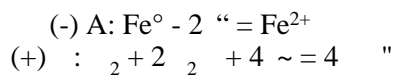
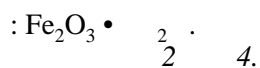
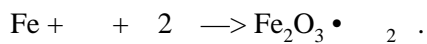
1 —

1, 2 3.

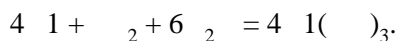
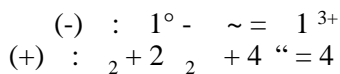
4,

5

() .



() :



()

()
 300³
 [2].

(HCl, H₂SO₄)
 Ni²⁺, ²⁺, ²⁺, ^{1³⁺} Meⁿ⁺ (Me — Cu²⁺,)
 ; ()

1.

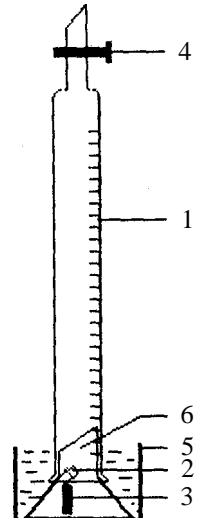
() [1].
 (-ff_{Ma}cc.)

() (S)
 :
 ^ . = ^7 (/ ' • S) -
 ()
 ()

2.

$$= \frac{h}{X} (/)$$

— / ³; 365 —
 ^ -365



- 1 —
- 2 —
- 3 —
- 4 —
- 5 —
- 6 —

3. (7), (5)

() (3) (4). 10

(2) 4. 1.

1. 3.

1—

			(. .), >1(),	/,	,	\wedge / 2	/,	
1								
2								
...								
1								
2								
...								
1								
2								
...								

5. () : () .

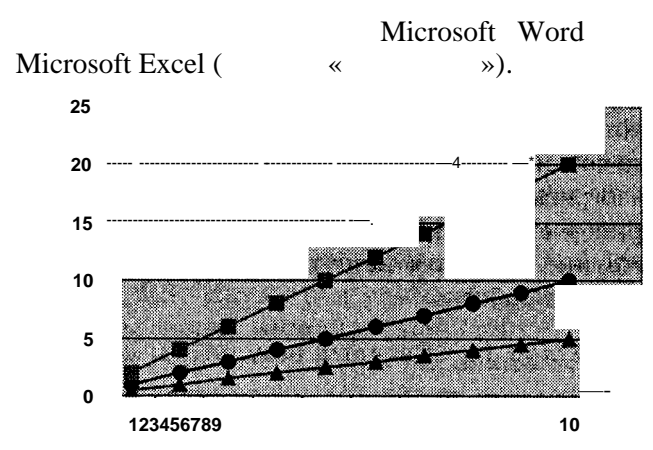
(. .). () : = . - 2 >

2 — , ; (. 2).

t, °C	15	16	17	18	19	20
, . . .	12,79	13,63	14,53	15,48	16,48	17,53
t, °C	15	16	17	18	19	20
, . . .	18,65	19,83	21,07	22,38	23,76	25,21

;
 (. . . 3).

$$t = \frac{273(\dots) - \dots}{(273 + t)760}$$
 ;
 10 (/)
 10 (0,005³)



Fe + 2HCl = FeCl₂ + H₂

$$= \frac{0,005 \cdot 56}{22,4} = 0,0125$$
 ;
 (p_{Zn} = 7190 /³);
 p_{Fe} — 7800 /³);

():

$$A = \frac{V_0}{V_1}$$
 ;
 6.

1. , 2002. — 71 /
2. « », 2012. — 50 /